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Covalent Bonding and Shapes of. Chapter (c) C-Cl $\hat{\text{I}} + \hat{\text{I}}\text{C-Cl}$
Chlorine is more electronegative than carbon Problem 1.6 Draw Lewis structures, showing all valence electrons, for these molecules.

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Covalent Bonding and Shapes of Molecules 1 CHAPTER 1

Solutions to the Problems Problem 1.1 Write and compare the ground-state electron configurations for each pair of elements: (a)

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Carbon and silicon C (6 electrons) $1s^2 2s^2 2p^2$ Si (14 electrons) $1s^2 2s^2 2p^6 3s^2 3p^2$ Both carbon and silicon have four electrons in their outermost (valence) shells.

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